

Assessment

Gases**Section Quiz: Gases and Pressure**

In the space provided, write the letter of the term or phrase that best completes each sentence or best answers each question.

- _____ 1. What *causes* a gas to exert pressure?
- collisions
 - density
 - temperature
 - elevation
- _____ 2. The SI unit for pressure is
- newton.
 - mm Hg.
 - pascal.
 - liter.
- _____ 3. The pressure exerted by a gas does *not* depend on
- temperature.
 - volume.
 - number of moles present.
 - the identity of the gas.
- _____ 4. At sea level, the average height of mercury in a barometer is
- 760 mm.
 - 101 325 atm.
 - 1.01 325 Pa.
 - All of the above
- _____ 5. Standard temperature and pressure are
- 32°F and 10 atm.
 - 0°C and 1 atm.
 - 10 K and 1 atm.
 - 0°F and 1 atm.

Section Quiz, *continued*

- _____ 6. Which of the following is *not* a unit of pressure?
- torr
 - pascal
 - newton
 - atmosphere
- _____ 7. A pressure of 760.0 mm Hg is equal to
- 1.000 atm.
 - 10.00 atm.
 - 100.0 atm.
 - 7.600 atm.
- _____ 8. A pressure of 20 torr is equal to
- 133.322 Pa.
 - 20 Pa.
 - 20 mm Hg.
 - 1 mm Hg.
- _____ 9. Who developed the concept that the total pressure of a mixture of gases is the sum of their partial pressures?
- Charles
 - Boyle
 - Dalton
 - Kelvin
- _____ 10. Gases collected by water displacement contain
- $\text{H}_2\text{O}(g)$.
 - $\text{H}_2\text{O}(l)$.
 - $\text{H}_2\text{O}(s)$ and $\text{H}_2\text{O}(l)$.
 - $\text{H}_2\text{O}(l)$ and $\text{H}_2\text{O}(g)$.