

Chemistry Nomenclature Review Name: \_\_\_\_\_

For the following compounds, give the formulas and the molar masses:

- |     | Formula                   | Molar Mass |
|-----|---------------------------|------------|
| 1)  | sodium phosphide _____    | _____      |
| 2)  | magnesium nitrate _____   | _____      |
| 3)  | lead (II) sulfite _____   | _____      |
| 4)  | calcium phosphate _____   | _____      |
| 5)  | ammonium sulfate _____    | _____      |
| 6)  | silver cyanide _____      | _____      |
| 7)  | aluminum sulfide _____    | _____      |
| 8)  | beryllium chloride _____  | _____      |
| 9)  | copper (I) arsenide _____ | _____      |
| 10) | iron (III) oxide _____    | _____      |

For the following compounds, identify the oxidation number of each atom.

- 11)  $\text{NH}_3$  \_\_\_\_\_
- 12)  $\text{HNO}_3$  \_\_\_\_\_
- 13)  $\text{SO}_4^{-2}$  \_\_\_\_\_
- 14)  $\text{SnF}_2$  \_\_\_\_\_
- 15)  $\text{MoO}_4^{-2}$  \_\_\_\_\_

Complete the following calculations. Show all work and use appropriate significant figures.

1) How many grams are in 4.5 moles of lithium oxide?

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2) How many atoms are in 23 moles of sodium bromide?

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3) How many moles are in  $3.4 \times 10^{23}$  atoms of barium nitrate?

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4) How many atoms are in 25 grams of ammonium sulfate?

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5) Identify the percentage composition of each element in glucose.

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