Chemistry Name:

Electron Configurations

 Without using a periodic table, identify the information regarding the element whose electron configuration is shown.

1. $\left[Ne\right]3s^{1}$
	1. Period
	2. Group
	3. Block
	4. Valence electrons
	5. Reactivity
	6. Metal, nonmetal, metalloid
2. $\left[Kr\right]5s^{2}4d^{10}5p^{5}$
	1. Period
	2. Group
	3. Block
	4. Valence electrons
	5. Reactivity
	6. Metal, nonmetal, metalloid
3. $\left[Ar\right]4s^{2}3d^{4}$
	1. Period
	2. Group
	3. Block
	4. Valence electrons
	5. Reactivity
	6. Metal, nonmetal, metalloid
4. $\left[Rn\right]7s^{2}$
	1. Period
	2. Group
	3. Block
	4. Valence electrons
	5. Reactivity
	6. Metal, nonmetal, metalloid
5. $\left[Ne\right]3s^{2}3p^{4}$
	1. Period
	2. Group
	3. Block
	4. Valence electrons
	5. Reactivity
	6. Metal, nonmetal, metalloid
6. $\left[Rn\right]7s^{2}6d^{1}5f^{3}$
	1. Period
	2. Block
	3. Valence electrons
	4. Reactivity
	5. Metal, nonmetal, metalloid
7. Write the electron configuration for the element in period 3 and group 2. Then identify the given information.
	1. Electron configuration
	2. Block
	3. Valence electrons
	4. Reactivity
	5. Metal, nonmetal, metalloid
8. Write the electron configuration for the element in period 5 with 7 valence electrons. Then identify the given information.
	1. Electron configuration
	2. Group
	3. Block
	4. Reactivity
	5. Metal, nonmetal, metalloid
9. Write the electron configuration for the element in period 4 and group 5. Then identify the given information.
	1. Electron configuration
	2. Block
	3. Valence electrons
	4. Reactivity
	5. Metal, nonmetal, metalloid